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#### **Key indicators**

Single-crystal X-ray study T = 292 K Mean  $\sigma$ (C–C) = 0.003 Å R factor = 0.047 wR factor = 0.119 Data-to-parameter ratio = 16.9

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

# 4-[(2-Chloro-5-methylphenoxy)acetoxymethyl]-2,6,7-trioxa-1-phosphabicyclo[2.2.2]octane 1-oxide

In the structure of the title compound,  $C_{14}H_{16}ClO_7P$ , the P atom is in a distorted tetrahedral configuration. The data reveal that some strain is probably present in the bicyclic structure. The terminal O=P bond distance is 1.4426 (16) Å, and the bridging P-O distances average 1.5728 (14) Å. The average value of the O=P-O angles is 114.36 (2)°, while the average value of the O-P-O angles is 104.18 (4)°.

# Comment

Heterocyclic compounds containing a symmetric caged bicyclic phosphate have received much attention since they were first synthesized (Verkade & Reynolds, 1960). Some of these molecules exhibit good biological activity, being particularly useful as herbicides (Ratz, 1966) and as flame retardants (Li *et al.*, 2002). The title compound, (I) (Fig. 1), has been prepared as part of our work on the synthesis of 2,6,7-trioxa-1-phosphabicyclo[2.2.2]octane-4-methanol 1-oxide and its alkoxylated derivatives and chloroacetate esters.





Selected bond lengths and angles are listed in Table 1. The O=P bond length is shorter than the P–O bond lengths within the bicyclic cage, as observed previously in a similar compound (Nimrod *et al.*, 1968). The external O=P–O angles are larger than the internal O–P–O angles, indicating a distorted tetrahedral configuration for the P atom. Atom C4, at the opposite end of the cage, exhibits a fairly normal tetrahedral geometry with the C–C–C angles ranging from 107.48 (16) to 111.51 (15)°. The bicyclic cage consists of three four-atom planes separated by dihedral angles of approximately 120°, as was noted earlier for a similar compound (Miu *et al.*, 1991).

# Experimental

2,6,7-Trioxa-1-phosphabicyclo[2.2.2]octane-4-methanol 1-oxide, (II), was prepared as described in the literature (Vyverberg & Chapman, 2002) in 95% yield. (2-Chloro-5-methylphenoxy)acetyl chloride, (III), was synthesized according to a literature method (Coutrot, 1986) in about theoretical yield. To a stirred solution of (II) (0.005 mol) and triethylamine (0.006 mol) in acetonitrile (25 ml), a solution of (III) (0.005 mol) in acetonitrile (5 ml) was added dropwise at 273–283 K. The mixture was then stirred at room temperature for

© 2006 International Union of Crystallography All rights reserved about 3 h. The solvent was then removed under reduced pressure and the residue was washed with water (20 ml). The raw product was recrystallized from acetonitrile, giving colorless block-shaped crystals of the title compound after 3 d.

Z = 4

 $D_{\rm x} = 1.547 {\rm Mg m}^{-3}$ 

Mo  $K\alpha$  radiation  $\mu = 0.38 \text{ mm}^{-1}$ 

T = 292 (2) K

 $R_{\rm int} = 0.071$ 

 $\theta_{\rm max} = 27.5^{\circ}$ 

Block, colorless

 $0.20 \times 0.20 \times 0.10 \text{ mm}$ 

3530 independent reflections

2528 reflections with  $I > 2\sigma(I)$ 

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.0648P)^2]$ 

where  $P = (F_0^2 + 2F_c^2)/3$ 

 $(\Delta/\sigma)_{\rm max} = 0.001$ 

 $\Delta \rho_{\rm max} = 0.27 \ {\rm e} \ {\rm \AA}^{-3}$ 

 $\Delta \rho_{\rm min} = -0.35 \text{ e} \text{ Å}^{-3}$ 

#### Crystal data

 $\begin{array}{l} C_{14}H_{16}ClO_7P\\ M_r = 362.69\\ Monoclinic, P2_1/n\\ a = 12.4504 \ (12) \ \text{\AA}\\ b = 6.3281 \ (6) \ \text{\AA}\\ c = 20.201 \ (2) \ \text{\AA}\\ \beta = 101.926 \ (2)^{\circ}\\ V = 1557.2 \ (3) \ \text{\AA}^3 \end{array}$ 

#### Data collection

Bruker SMART CCD area-detector diffractometer  $\varphi$  and  $\omega$  scans Absorption correction: none 9387 measured reflections

### Refinement

Refinement on  $F^2$   $R[F^2 > 2\sigma(F^2)] = 0.047$   $wR(F^2) = 0.119$  S = 0.963530 reflections 209 parameters

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Table 1
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Selected geometric parameters (Å, °).

P1-01	1.4426 (16)	P1-O4	1.5724 (14)
P1-O3	1.5696 (14)	P1-O2	1.5766 (15)
O1 B1 O2	114.01 (0)	C5 C4 C3	110.04 (15)
01 - P1 - 03 01 - P1 - 04	114.01(9) 114.03(9)	$C_{5}=C_{4}=C_{5}$	110.94 (13)
O3-P1-O4	104.15 (7)	C3-C4-C2	109.37 (16)
O1-P1-O2	115.02 (9)	C5-C4-C1	108.72 (15)
O3-P1-O2	103.99 (8)	C3-C4-C1	107.48 (16)
O4-P1-O2	104.39 (8)	C2-C4-C1	108.70 (16)
C4-C1-O2-P1	1.4 (2)		

H atoms were refined using a riding model, with C-H = 0.93 Å and  $U_{iso}(H) = 1.2U_{eq}(C)$  for aromatic, C-H = 0.98 Å and  $U_{iso}(H) = 1.2U_{eq}(C)$  for CH and C-H = 0.96 Å and  $U_{iso}(H) = 1.5U_{eq}(C)$  for methyl H atoms.

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINT* (Bruker, 2000); data reduction: *SAINT*; program(s) used to solve



#### Figure 1

The molecular structure of (I), showing 50% probability displacement ellipsolids and the atom-labeling scheme.

structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXTL* (Bruker, 1997); software used to prepare material for publication: *SHELXTL*.

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## References

Bruker (1997). SHELXTL. Version 5.10. Bruker AXS Inc., Madison, Wisconsin, USA.

- Bruker (2000). *SMART* (Version 5.059) and *SAINT* (Version 6.01). Bruker AXS Inc., Madison, Wisconsin, USA.
- Coutrot, P. (1986). Synthesis, pp. 661-664.
- Li, X., Ou, Y. X. & Zhang, Y. H. (2002). Chem. J. Chin. Univ. 23, 695-699.
- Miu, F. M., Liu, X. L., Li, Y. G., Wang, J. J., Liu, Y. S., Bao, J. C., Cao, J. H. & Zhou, W. (1991). Acta Chim. Sinica, 49, 870–875.
- Nimrod, D. M., Fitzwater, D. R. & Verkade, J. G. (1968). J. Am. Chem. Soc. 90, 2780–2784.
- Ratz, R. F. W. (1966). US Patent No. 3 287 448.
- Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.
- Verkade, J. G. & Reynolds, L. J. (1960). J. Org. Chem. 25, 663-667.
- Vyverberg, F. J. & Chapman, R. W. (2002). US Patent No. 6 455 722.